Adult Education Class Outline
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1.2031 Physical Science

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Title:	Physical Science
Submitted by:	Yosemite Adult Education

Yosemite Joint Union High School District
Yosemite Adult Education
Course Outline

Course Title: Physical Science
Program: Adult Independent Study
State Course Code: 1.2031
Date Submitted: 01/09/06
Grade Level: 9-12
Length of Course: 1 year
Credits: 10

Textbook: *Physical Science,* Glencoe, 2005
Supplemental Material: *Study Guide and Reinforcement, Student Edition—Glencoe, 2005*

I. Course Description; This course is designed to help students learn the properties of matter, elements, compounds, electricity, sound, and light. It will also increase science literacy, improve reading comprehension, and deepen students' understanding of ideas and concepts.

II. Major Goals and Purpose of the Course: The student will learn about all areas of physical science. They will be able to use scientific skills to solve problems and answer questions in all the areas. The student will also discover how important physical Science is to their life.

III. Course outline/Alignment to California State Standards and Expected Schoolwide Learning Results:

Chapter/ Text	Supplemental Material	Assignment Course Content/Objective	Standards Addressed	ESLRS
1	1-3	The Nature of Science	2.4e, 3.1a, 3.1c, 3.1f, 3.1g, 3.1h, 3.15	1,2,3,4,6 2

5-7	Motion		1.1a, 1.1b, 1.1h, 1.1j, 1.1k, 3.1a, 3.1h, 3.1n	1,2,3,4,6	3 4
13-14 9-11	Energy Forces		1.2a-1.2c, 1.2h, 1.1c-1.1m, 1.3c, 1.5o, 1.2e-1.2g, 3.1e, 3.1n	1,2,3,4,6 1,2,3,4,6	5 6
15-17	Work & Machines		1.3a, 3.1b	1,2,3,4,6	7
19-21	Thermal Energy		1.3a, 1.3b, 1.3c, 1.3g, 2.7a, 2.7d, 3.1b, 3.1l	1,2,3,4,6 23-25	8 9
Electricity		1.1m, 1.5a, 1.5b, 1.5c, 1.5e, 1.5j, 2.7c		1,2,3,4,6	27-29
Magnetism & Its Uses					10
1.5f, 1.5g, 1.5h, 1.5j, 3.1i	1,2,3,4,6	31-33	Energy Sources		11
2.6f, 3.1m	1,2,3,4,6	35-37	Waves		12
1.4a, 1.4b, 1.4c, 1.4d, 1.4f	1,2,3,4,6	39-42	Sound		13
1.4b, 1.4d, 1.4f, 2.3g	1,2,3,4,6	43-45	Electromagnetic Waves		14
1.4a, 1.4e, 3.1n	1,2,3,4,6	47-50	Light		15
1.4f	1,2,3,4,6	51-53	Mirrors and Lenses		16
	1,2,3,4,6	55-56	Classification of Matter		17
2.6f	1,2,3,4,6	57-59	Solids, Liquids & Gases		18
1.3c, 1.5i, 2.4g,	1,2,3,4,6				

2.2d, 2.2h, 2.4a		61-63	Properties of Atoms & The Periodic Table	19	
				20	
2.1a, 2.1b, 2.1c, 2.1d, 2.1d, 3.1i, 2.1e, 2.1g, 2.2e, 2.11g	1,2,3,4,6	65-68	Radioactivity & Nuclear Reactions of Multiplication		
		f, 2.11a, 2.11b-2.11f, 3.1e	1,2,3,4,6	69-71	1,2,3,4,6
Elements & Their Properties		2.1c, 2.3b, 2.3d, 2.3e, 2.3f, 2.2g, 2.2h	73-75	2.1d, 2.2a, 2.2b, 2.1c, 2.2e, 2.1d, 2.2a, 2.2b, 2.1c, 2.2e,	
Chemical Bonds		1,2,3,4,6	21	77-80	2.3a, 2.7b, 2.8c, 2.3e, 2.3f, 2.4d
			Chemical Reactions		1,2,3,4,6
22	81-84	Solutions		2.6a-2.6e, 2.4h, 2.4i, 2.5e, 2.5f	1,2,3,4,6
23	85-87	Acids, Bases & Solids		2.3c, 2.5a-2.5g, 2.7e, 2.7f	1,2,3,4,6
24	89-92	Organic Compounds		2.2b, 2.10a, 2.10b-2.10f, 3.1n	1,2,3,4,6
25	93-95	New Materials Through Chemistry		3.1c, 3.1g, 2.7f, 2.8a-2.8d, 2.9a-2.9c	1,2,3,4,6

Science Content Standards.

Standards that all students are expected to achieve in the course of their studies are unmarked. Standards that all students should have the opportunity to learn are marked with an asterisk (*).
Atomic and Molecular Structure

1. The periodic table displays the elements in increasing atomic number and shows how periodicity of the physical and chemical properties of the elements relates to atomic structure. As a basis for understanding this concept:
    1. Students know how to relate the position of an element in the periodic table to its atomic number and atomic mass.
    2. Students know how to use the periodic table to identify metals, semimetals, nonmetals, and halogens.
    3. Students know how to use the periodic table to identify alkali metals, alkaline earth metals and transition metals, trends in ionization energy, electronegativity, and the relative sizes of ions and atoms.
    4. Students know how to use the periodic table to determine the number of electrons available for bonding.

5. Students know the nucleus of the atom is much smaller than the atom yet contains most of its mass.
6. * Students know how to use the periodic table to identify the lanthanide, actinide, and transactinide elements and know that the transuranium elements were synthesized and identified in laboratory experiments through the use of nuclear accelerators.
7. * Students know how to relate the position of an element in the periodic table to its quantum electron configuration and to its reactivity with other elements in the table.
8. * Students know the experimental basis for Thomson's discovery of the electron, Rutherford's nuclear atom, Millikan's oil drop experiment, and Einstein's explanation of the photoelectric effect.
9. * Students know the experimental basis for the development of the quantum theory of atomic structure and the historical importance of the Bohr model of the atom.
   * Students know that spectral lines are the result of transitions of electrons between energy levels and that these lines correspond to photons with a frequency related to the energy spacing between levels by using Planck's relationship (E = hv).

Chemical Bonds

2. Biological, chemical, and physical properties of matter result from the ability of atoms to form bonds from electrostatic forces between electrons and protons and between atoms and molecules. As a basis for understanding this concept:
   1. Students know atoms combine to form molecules by sharing electrons to form covalent or metallic bonds or by exchanging electrons to form ionic bonds.
   2. Students know chemical bonds between atoms in molecules such as H2 , CH4 , NH3 , H2 CCH2 , N2 , Cl2 , and many large biological molecules are covalent.
   3. Students know salt crystals, such as NaCl, are repeating patterns of positive and negative ions held together by electrostatic attraction.
   4. Students know the atoms and molecules in liquids move in a random pattern relative to one another because the intermolecular forces are too weak to hold the atoms or molecules in a solid form.
   5. Students know how to draw Lewis dot structures.
   6. * Students know how to predict the shape of simple molecules and their polarity from Lewis dot structures.
   7. * Students know how electronegativity and ionization energy relate to bond formation.
   8. * Students know how to identify solids and liquids held together by van der Waals forces or hydrogen bonding and relate these forces to volatility and boiling/ melting point temperatures.

Conservation of Matter and Stoichiometry

3. The conservation of atoms in chemical reactions leads to the principle of conservation of matter and the ability to calculate the mass of products and reactants. As a basis for understanding this concept:
   1. Students know how to describe chemical reactions by writing balanced equations.
   2. Students know the quantity one mole is set by defining one mole of carbon 12 atoms to have a mass of exactly 12 grams.
   3. Students know one mole equals 6.02x1023particles (atoms or molecules).

4. Students know how to determine the molar mass of a molecule from its chemical formula and a table of atomic masses and how to convert the mass of a molecular substance to moles, number of particles, or volume of gas at standard temperature and pressure.
5. Students know how to calculate the masses of reactants and products in a chemical reaction from the mass of one of the reactants or products and the relevant atomic masses.
6. * Students know how to calculate percent yield in a chemical reaction.
7. * Students know how to identify reactions that involve oxidation and reduction and how to balance oxidation-reduction reactions.

Gases and Their Properties

4. The kinetic molecular theory describes the motion of atoms and molecules and explains the properties of gases. As a basis for understanding this concept:
   Students know the random motion of molecules and their collisions with a surface create the observable pressure on that surface.
   1. Students know the random motion of molecules explains the diffusion of gases.
   2. Students know how to apply the gas laws to relations between the pressure, temperature, and volume of any amount of an ideal gas or any mixture of ideal gases.
   3. Students know the values and meanings of standard temperature and pressure (STP).
   4. Students know how to convert between the Celsius and Kelvin temperature scales.
   5. Students know there is no temperature lower than 0 Kelvin.
   6. * Students know the kinetic theory of gases relates the absolute temperature of a gas to the average kinetic energy of its molecules or atoms.
   7. * Students know how to solve problems by using the ideal gas law in the form PV = nRT.
   8. * Students know how to apply Dalton's law of partial pressures to describe the composition of gases and Graham's law to predict diffusion of gases.

Acids and Bases

5. Acids, bases, and salts are three classes of compounds that form ions in water solutions. As a basis for understanding this concept:
   1. Students know the observable properties of acids, bases, and salt solutions.
   2. Students know acids are hydrogen-ion-donating and bases are hydrogen-ion-accepting substances.
   3. Students know strong acids and bases fully dissociate and weak acids and bases partially dissociate.
   4. Students know how to use the pH scale to characterize acid and base solutions.
   5. * Students know the Arrhenius, Brønsted-Lowry, and Lewis acid-base definitions.
   6. * Students know how to calculate pH from the hydrogen-ion concentration.
   7. * Students know buffers stabilize pH in acid-base reactions.

Solutions

6. Solutions are homogeneous mixtures of two or more substances. As a basis for understanding this concept:

1. Students know the definitions of solute and solvent.
2. Students know how to describe the dissolving process at the molecular level by using the concept of random molecular motion.
3. Students know temperature, pressure, and surface area affect the dissolving process.
4. Students know how to calculate the concentration of a solute in terms of grams per liter, molarity, parts per million, and percent composition.
5. * Students know the relationship between the molality of a solute in a solution and the solution's depressed freezing point or elevated boiling point.
6. * Students know how molecules in a solution are separated or purified by the methods of chromatography and distillation.

Chemical Thermodynamics

7. Energy is exchanged or transformed in all chemical reactions and physical changes of matter. As a basis for understanding this concept:
    1. Students know how to describe temperature and heat flow in terms of the motion of molecules (or atoms).
    2. Students know chemical processes can either release (exothermic) or absorb (endothermic) thermal energy.
    3. Students know energy is released when a material condenses or freezes and is absorbed when a material evaporates or melts.
    4. Students know how to solve problems involving heat flow and temperature changes, using known values of specific heat and latent heat of phase change.
    5. * Students know how to apply Hess's law to calculate enthalpy change in a reaction.
    6. * Students know how to use the Gibbs free energy equation to determine whether a reaction would be spontaneous.

Reaction Rates

8. Chemical reaction rates depend on factors that influence the frequency of collision of reactant molecules. As a basis for understanding this concept:
    1. Students know the rate of reaction is the decrease in concentration of reactants or the increase in concentration of products with time.
    2. Students know how reaction rates depend on such factors as concentration, temperature, and pressure.
    3. Students know the role a catalyst plays in increasing the reaction rate.
    4. * Students know the definition and role of activation energy in a chemical reaction.

Chemical Equilibrium

9. Chemical equilibrium is a dynamic process at the molecular level. As a basis for understanding this concept:
    1. Students know how to use Le Chatelier's principle to predict the effect of changes in concentration, temperature, and pressure.
    2. Students know equilibrium is established when forward and reverse reaction rates are equal.
    3. * Students know how to write and calculate an equilibrium constant expression for a reaction.

Organic Chemistry and Biochemistry

10. The bonding characteristics of carbon allow the formation of many different organic molecules of varied sizes, shapes, and chemical properties and provide the biochemical basis of life. As a basis for understanding this concept:
    1. Students know large molecules (polymers), such as proteins, nucleic acids, and starch, are formed by repetitive combinations of simple subunits.
    2. Students know the bonding characteristics of carbon that result in the formation of a large variety of structures ranging from simple hydrocarbons to complex polymers and biological molecules.
    3. Students know amino acids are the building blocks of proteins.
    4. * Students know the system for naming the ten simplest linear hydrocarbons and isomers that contain single bonds, simple hydrocarbons with double and triple bonds, and simple molecules that contain a benzene ring.
    5. * Students know how to identify the functional groups that form the basis of alcohols, ketones, ethers, amines, esters, aldehydes, and organic acids.
    6. * Students know the R-group structure of amino acids and know how they combine to form the polypeptide backbone structure of proteins.

Nuclear Processes

11. Nuclear processes are those in which an atomic nucleus changes, including radioactive decay of naturally occurring and human-made isotopes, nuclear fission, and nuclear fusion. As a basis for understanding this concept:
    1. Students know protons and neutrons in the nucleus are held together by nuclear forces that overcome the electromagnetic repulsion between the protons.
    2. Students know the energy release per gram of material is much larger in nuclear fusion or fission reactions than in chemical reactions. The change in mass (calculated by E = mc2) is small but significant in nuclear reactions.
    3. Students know some naturally occurring isotopes of elements are radioactive, as are isotopes formed in nuclear reactions.
    4. Students know the three most common forms of radioactive decay (alpha, beta, and gamma) and know how the nucleus changes in each type of decay.
    5. Students know alpha, beta, and gamma radiation produce different amounts and kinds of damage in matter and have different penetrations.
    6. * Students know how to calculate the amount of a radioactive substance remaining after an integral number of half-lives have passed.
    7. * Students know protons and neutrons have substructures and consist of particles called quarks.

V. Outcomes:
Students will demonstrate basic competency in course standards by achieving a minimum 70% proficiency on coursework or by passing the final exam.

VI. Instructional Strategies:
Direct instruction
Individual instruction
Peer tutoring

VII. Hours of Instruction: 75 hours per semester

VIII. Evaluation:

Publisher prepared tests and quizzes
Review of student work samples
Student demonstrations
Other informal assessments.

IX. Repetition:
After a student has successfully completed this course, he/she may not be allowed to repeat course.